Electrons

Fill out how many protons, neutrons, and electrons each atom has using the information given. Then draw the electrons in each shell, remembering that the first shell can hold 2 electrons, the second can hold 8 electrons, and the third can hold 18 electrons. Finally, answer the questions at the bottom.

Atomic No: 2
Mass No: 4
Protons: 2
Neutrons: 2
Electrons: 2

Atomic No: 10
Mass No: 20
Protons: 10
Neutrons: 10
Electrons: 10

Atomic No: 1
Mass No: 1
Protons: 1
Neutrons: 1
Electrons: 1

Atomic No: 9
Mass No: 19
Protons: 9
Neutrons: 9
Electrons: 9

Atomic No: 11
Mass No: 23
Protons: 11
Neutrons: 11
Electrons: 11

Which elements would be most likely to lose electrons in a chemical bond? ______

Which elements would be most likely to gain electrons in a chemical bond? ______
Electrons

Fill out how many protons, neutrons, and electrons each atom has using the information given. Then draw the electrons in each shell, remembering that the first shell can hold 2 electrons, the second can hold 8 electrons, and the third can hold 18 electrons. Finally, answer the questions at the bottom.

Atomic No: 2
Mass No: 4
Protons: 2
Neutrons: 2
Electrons: 2

Atomic No: 10
Mass No: 20
Protons: 10
Neutrons: 10
Electrons: 10

Atomic No: 1
Mass No: 1
Protons: 1
Neutrons: 0
Electrons: 1

Atomic No: 9
Mass No: 19
Protons: 9
Neutrons: 10
Electrons: 9

Atomic No: 11
Mass No: 23
Protons: 11
Neutrons: 12
Electrons: 11

Which element(s) would be most likely to lose electrons in a chemical bond? **H, Na**
(both only have one electron in their valence shell)

Which element(s) would be most likely to gain electrons in a chemical bond? **H, F**
(both are only missing one electron in their valence shell)